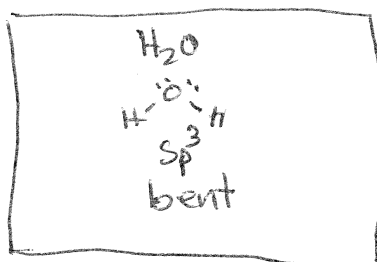


On the white boards: (1) Draw structure
(2) hybridization, (3) shape.

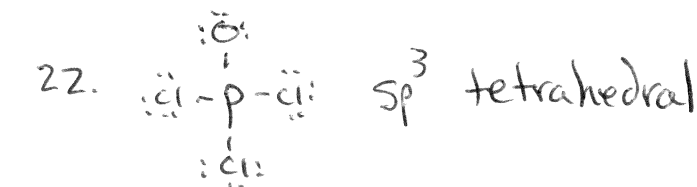
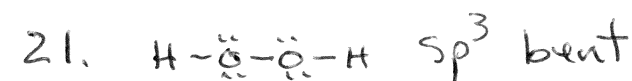
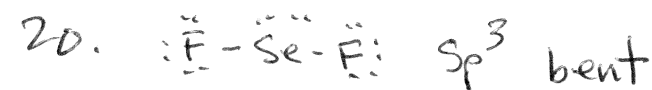
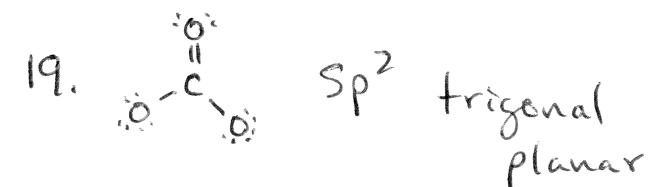
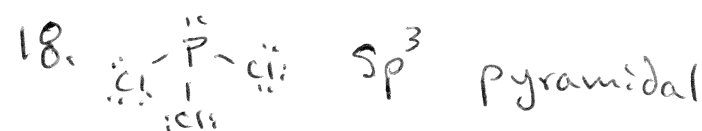
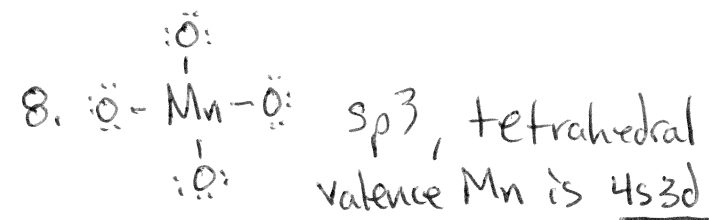
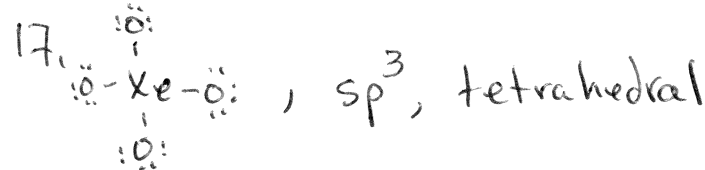
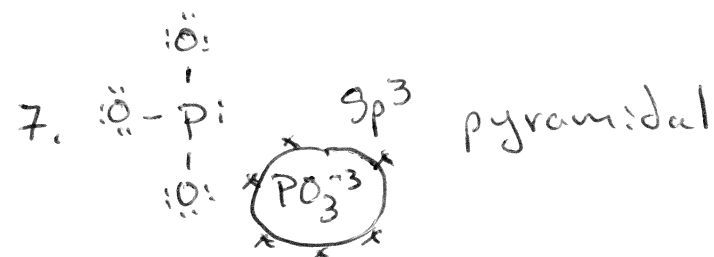
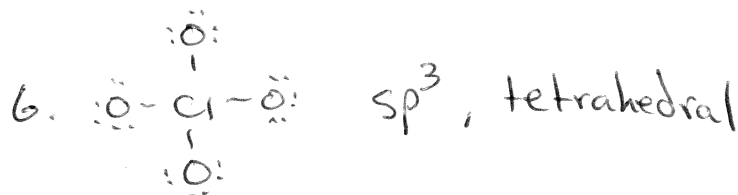
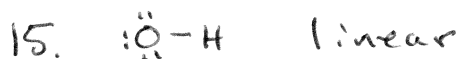
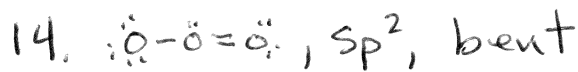
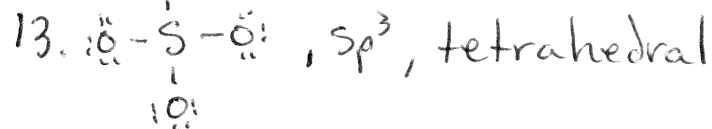
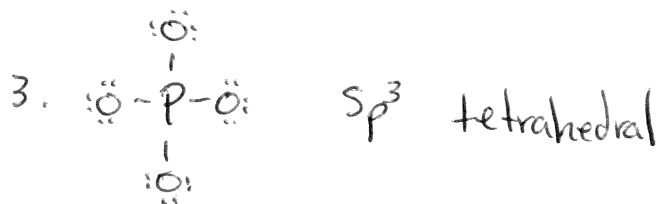
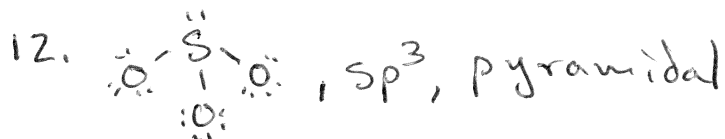
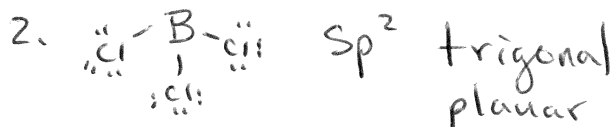
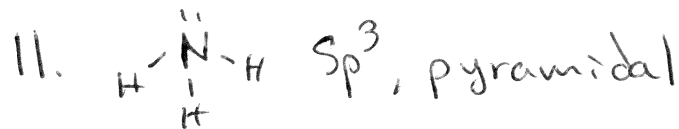
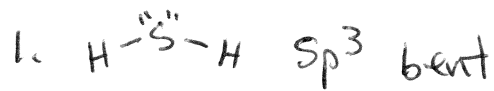
Example

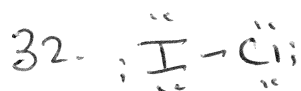
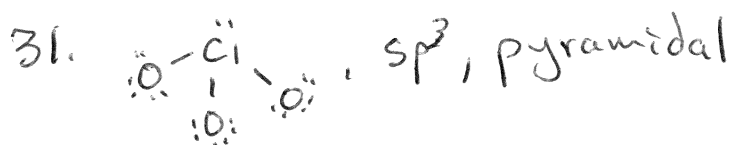
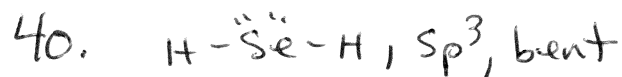
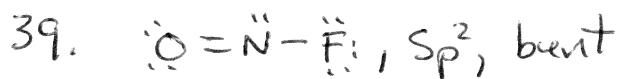
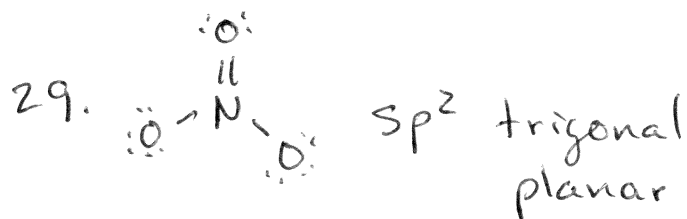
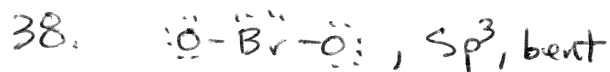
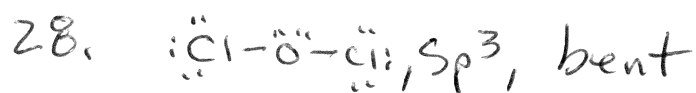
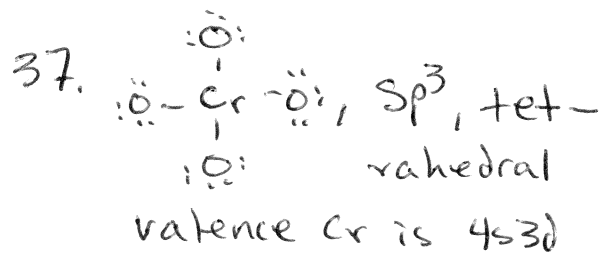
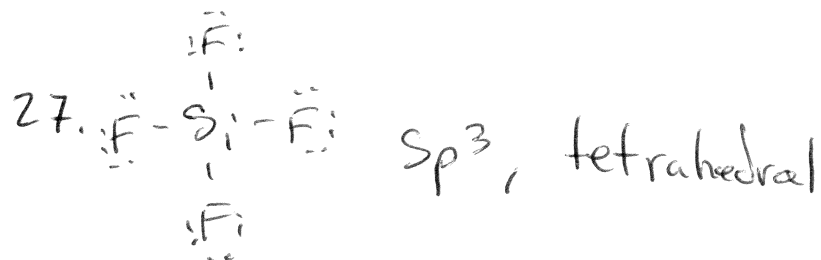
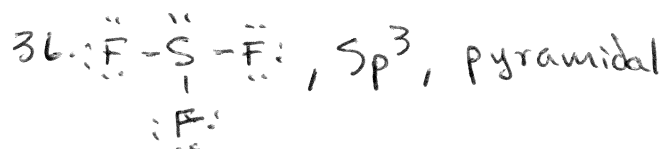
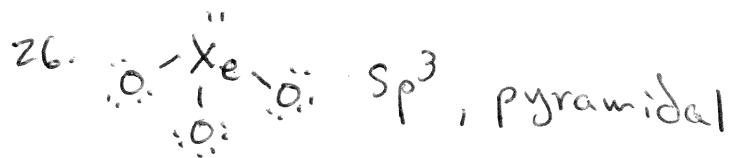
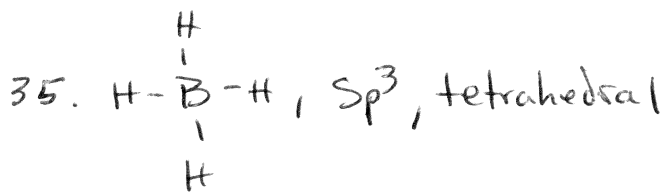
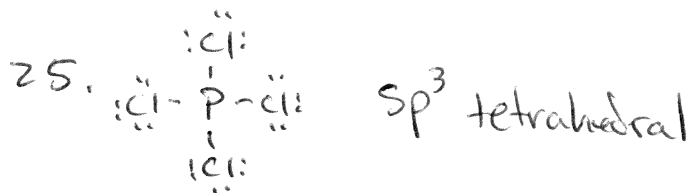
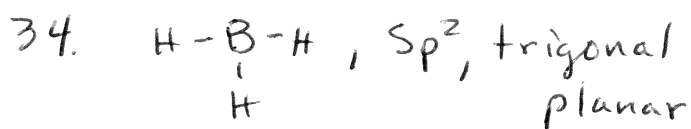
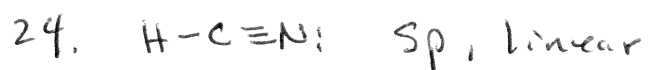
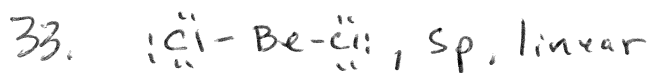


Key

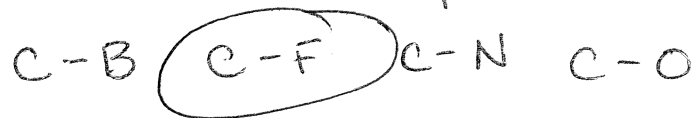
- | | | |
|-------------------------|----------------------------|-------------------------|
| 1. H_2S | 13. SO_4^{2-} | 29. NO_3^- |
| 2. BCl_3 | 14. O_3 | 30. NO_2^- |
| 3. PO_4^{3-} | 15. OH^- | 31. ClO_3^- |
| 4. ClO^- | 16. XeO_2 | 32. ICl |
| 5. ClO_2^- | 17. XeO_4 | 33. BeCl_2 |
| 6. ClO_4^- | 18. PCl_3 | 34. BH_3 |
| 7. PO_3^{2-} | 19. CO_3^{2-} | 35. BH_4^- |
| 8. MnO_4^- | 20. SeF_2 | 36. SF_3^+ |
| 9. CO_2 | 21. H_2O_2 | 37. CrO_4^{2-} |
| 10. NCl_3 | 22. POCl_3 | 38. BrO_2^- |
| 11. NH_3 | 23. CN^- | 39. ONF |
| 12. SO_3^{2-} | 24. HCN | 40. SeH_2 |
| | 25. PCl_4^+ | |
| | 26. XeO_3 | |
| | 27. SiF_4 | |
| | 28. OCl_2 | |

Lewis Formula Key for Test





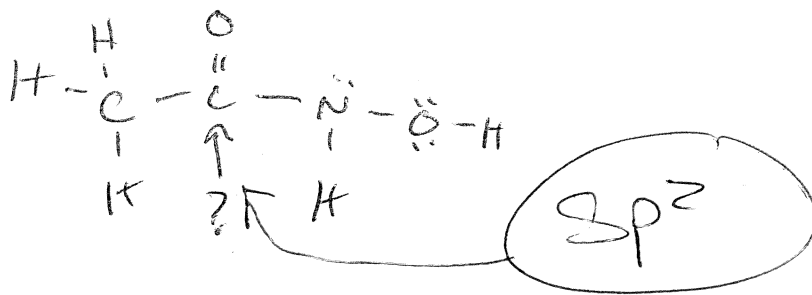
41. Select the most polar bond:



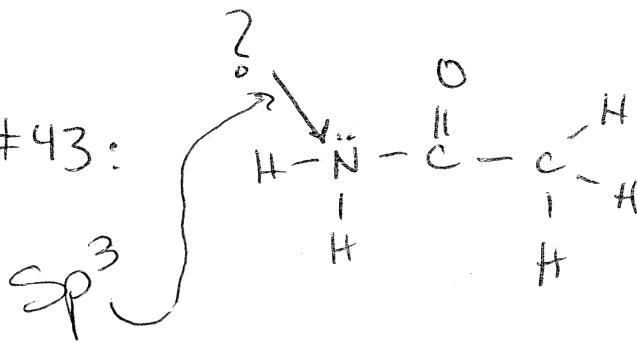
42. Select the most polar bond:



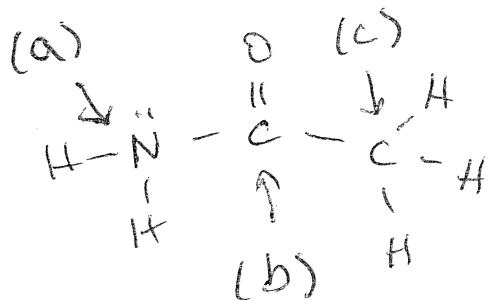
43. Determine hybridization at atom indicated by arrow:



44. Same as #43:



45. Identify geometry in order a, b, c:



a = pyramidal

b = trigonal planar

c = tetrahedral