

Physical Properties Compounds

<u>Element or compound type</u>	<u>bond</u>	(MP) <u>melting point</u>	<u>conductivity</u>	<u>Example</u>
Molecular Covalent	Intermolecular Force (IMF)	Low For H ₂ O mp = 0°C	Low No ions! e ⁻ cannot move.	H ₂ O
Ionic	Ionic Bond between ions	High For NaCl mp = 801°C	High when substance is <u>molten</u> or <u>melted</u> only.	NaCl
Metals	Sea of electrons shared among atoms	High For Al, mp = 660°C For Fe, mp = 1538°C	High e ⁻ move. <u>e⁻ sea</u>	Al, Fe

conductivity: ability to carry or support an electrical current. To carry a current electrons and ions must be free to move!